**Unit 6 Review Sheet**

*Your test will consist primarily of word equations for which you must write the* ***molecular, total ionic, and net ionic equations****. The examples listed below are comprehensive of everything we have studied in this unit*. *Write each of the three equations listed above for each reaction.*

1. A solution of cadmium (II) nitrate is added to a solution of cesium sulfide.

2. Aqueous sulfuric acid (H2SO4) is mixed with a solution of barium hydroxide.

3. A solution of hydrochloric acid (HCl) is added to solid iron (II) sulfide producing hydrosulfuric acid (H2S) and aqueous iron (II) chloride.

4. Solid calcium carbonate is added to aqueous hydrochloric acid (HCl) producing calcium chloride, water, and carbon dioxide gas.

5. Solutions of calcium nitrate and rubidium chloride are mixed.

6. Magnesium turnings are added to a solution of iron (III) chloride.

7. Sodium metal is added to water.

8. Lithium is added to hydrochloric acid (HCl)

9. Chlorine gas is bubbled through a solution of potassium iodide producing solid iodine and potassium chloride.

10. Zinc metal is added to a solution of sodium chloride.

11. A piece of copper is dropped into a container of water.

12. Zinc pellets are added to a sulfuric acid (H2SO4) solution.

13. A solution of iron (III) chloride is poured over a piece of platinum wire.

14. Magnesium turnings are added to a solution of lead (II) acetate.

15. Aqueous solutions of potassium bromide and silver (I) nitrate are mixed.